INDEX

A

Alkaloid - 153

Aqueous medium – 110,114,123

Allopurinol - 11, 89,124,125

Apparent molar volume - 49, 51,124,126

Association constant - 27, 33, 70, 240

Amino acid – 1, 2, 3, 36

B

Biochemic Body - 1,124,134, 267

Bioactive molecule - 9,266

Biomolecule - 2

Binding Constant - 24, 200, 201, 235

Benesi-Hildebrand method - 24,150, 200, 240

C

Crown ether - 93, 94, 110, 111

Complexation - 16, 19, 25, 26

Conductance, 69, 73,115, 209

Conductivity measurement - 23,135,213,161,190

Covalent bond - 4,13,14,15

Conductivity - 13, 23, 29, 36

α-cyclodextrin - 90,91,154,244

β-cyclodextrin - 10, 21,91, 302

D

2D ROESY – 9, 24, 26,199

1, 3-dimethylimidazolium methyl sulphate – 9,175

Density measurement – 9, 99, 175

Designer solvents – 13, 174

Dipole-dipole forces – 5, 34

Dipole-induced dipole force - 34

E

Electrolytes – 43, 44 ,50, 51

Electrostatic interaction – 44, 50,51

Enthalpy – 130, 132, 241, 242

Entropy – 33, 41, 48, 63

2-ethoxyethanol – 82,178

F

Falkenhagen-vernon equation - 58

Fluidity – 60,178

Formation Constant – 24,179

FTIR Spectroscopy – 26,104

Fuoss conductance equation – 29,177,188

G

Gibbs energy - 48,188

Green solvents - 13,174

H

HP-β-Cyclodextrin – 10,153,154, 267

H-bonding – 5,16,35,43

Hydrophilic – 20,145,148,198

Hydrophobic – 7, 8, 20, 22

Hydrophobic interactions – 22,33,152

I

Inclusion complex – 6,7,8,9,10

Industrial application - 82

Intermolecular force – 14,37,46,47

Intermolecular interaction – 37,46,47,52

Ion-association - 178,179,180,184

Ion-dipole force – 38,43

Ionic liquids – 8,9,10,12

Ionic mobility - 64

Ion-pair - 29, 30

J

Jones-dole equation - 58,60,61,128

K

Kohlrausch's equation - 31

Kohlrausch's law - 72

L

London Dispersion Force - 34, 35

Lattice energy – 33, 38

Limiting apparent molar volume – 9,126,129,130

Limiting equivalent conductance – 72, 73

M

Mass Measurement of mass - 96

Masson equation – 49,126,127

Melting point – 12,174

Metallic bonding - 37

Methanol –81

Ethanol – 175,176

2-methoxyethanol – 81,178

N

Nortriptyline hydrochloride – 89

Neoteric solvent – 85, 86

Non-aqueous electrolyte – 273,299

Non-covalent interactions-201, 266

NMR Spectroscopy – 25,152, 202, 238

O

Optical data – 74

Organic solvent – 6,13, 20, 42

P

2-propoxy ethanol – 176,178,179

Partial molar volumes – 5, 9 ,44, 48

Pharmaceutical – 68, 91, 92, 144

Physicochemical properties – 4,174, 235

Pitzer equation – 49, 51

R

Redlich-kister expression - 76

Reference electrolyte - 180

Refractive index – 24, 41, 74, 75

Relative permittivity – 177,178,185

Repulsive interaction - 70

Room temperature ionic liquid - 235

S

Solvation – 123,129,179,195

Solvation consequence – 123,195,269

Shedlovsky extrapolation - 177

Stokes' law – 53,73,179

Surface Tension – 10, 24, 29,104

T

Theophylline – 11,145,146,153

Thiamine hydrochloride – 3

Thermodynamic properties – 4,5,6,9

Thermostat water bath – 97,98

Thermodynamic parameters – 10,33,150,154

U

Uric Acid – 17,87,90,124

UV-Visible Spectra - 133,142,146,147

V

van der waal forces – 12,15,34,128

Viscosity measurement – 103,184, 267, 268

Viscosity – 6, 9, 10, 12

Viscous flow – 54,55,62,67

W

Water – 1,2,4,6