

List of Publications

1. Moazzam Ali, Mrinmoy Jha, Susanta K. Das and Swapan K. Saha. Hydrogen-Bond-Induced Microstructural Transition of Ionic Micelles in the Presence of Neutral Naphthols: pH Dependent Morphology and Location of Surface Activity. *J. Phys. Chem. B* **2009**, *113*, 15563–15571.
2. Swapan K. Saha, Mrinmoy Jha, Moazzam Ali, Amitabha Chakraborty, Goutam Bit, and Susanta K. Das. Micellar Shape Transition under Dilute Salt-Free Conditions: Promotion and Self-Fluorescence Monitoring of Stimuli-Responsive Viscoelasticity by 1- and 2-Naphthols. *J. Phys. Chem. B* **2008**, *112*, 4642-4647.
3. Moazzam Ali, Amitabha Chakraborty, Saumik Bardhan and Swapan K. Saha. Thermodynamics of Micellisation of Ammoniumdodecyl Sulfate in Aqueous Solutions of Symmetrical Tetraalkylammonium Bromides: Clouding in Presence of Tetrabutylammonium Bromide salt. *Journal of Dispersion Science and Technology* **2010**, *31*, 1220-1229.
4. Moazzam Ali, Soumik Bardhan and Swapan K. Saha. Polyoxyethylene Sorbitan Monolaurate (Tween-20) induced vesicle-to-micelle transition of aqueous dimethyldioctadecylammonium bromide dispersion. *Colloids Surf. A. (Communicated)*
5. Moazzam Ali and Swapan K. Saha. **Hydrogen Bonded Large Molecular Aggregates of Charged Amphiphiles and Unusual Rheology: Photochemistry and Photophysics of Hydroxyaromatic Dopants**. In “Hydrogen Bonding and Transfer in the Excited-State” Ed- Ke-Li Han and Guang-Jiu Zhao. Vol. II, Chapter 31, pages 711-745, John Wiley & Sons, Ltd. 2010.

Addendum and Corrigendum
for the Thesis entitled:

**Physico-Chemical Studies on Selected Amphiphiles and Their Aggregation
Behaviour in Different Media**

1. Reference related to Figure 1.3


Israelachvili, J.N. In Intermolecular and Surface Forces, Academic Press, London,
1994


Reference related to Table 1.1

Israelachvili, J.N.; Mitchel, D.J.; Ninham, B.W. *J. Chem. Soc., Faraday Trans. 2* 72,
1525, 1976.

2. The numerical values of β in Table 3.1 were derived from the slope of the lines in Figure 3.5.
3. The word 'figure', throughout the text, should be read as 'Figure'.
4. In section 3.3.2, the Corrin-Harkins equation as referred to, is usually meant for salt with a common ion, but recent studies have shown that such plots are also applicable to other electrolytes (Behera. K.; Pandey. S. *J. Colloid Interface Sci.* 316, 803-814, 2007.
5. The proposition in Figure 4.44 could not be justified by means of measured values of Γ because the probe molecules are undergoing interaction with the micellar aggregates and not with that of water surface.
6. The wavelength at which the absorption intensities of the probe molecules were measured for calculating the binding constant values with the surfactants are summarized below:

1-Naphthol	321.2 nm
2-Naphthol	327.3 nm
2,3-Dihydroxynaphthalene	323.9 nm
2,7-Dihydroxynaphthalene	325.7 nm


Prof. S. K. Saha 19.09.2011
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